**Atomic Mass**

* The **atomic number** represents the number of protons in the nucleus.
* Atomic Mass Unit (AMU)
	+ 1 Proton = 1 AMU
	+ 1 Electron =0.0006 AMU
	+ 1 Neutron = 1 AMU
* The number of protons must equal the number of electrons for the atom to be stable (meaning electrically neutral).
* An ion is an atom that is electrically charged.
	+ A charge is created when electrons are transferred to or from another atom.



* The **atomic mass** is the combined mass of the protons and neutrons in the nucleus. \*Also known as the “Mass Number”
* Atoms of the same element ***ALWAYS*** have the same number of protons, but the number of neutorns may vary.
* An isotope is an atom with the same number of protons and different number of neutrons than other atoms of that element.

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